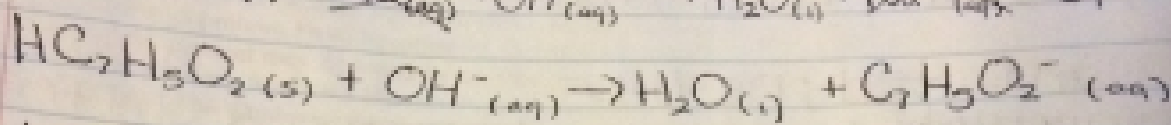
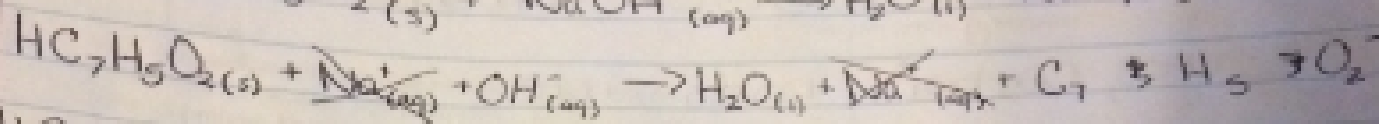
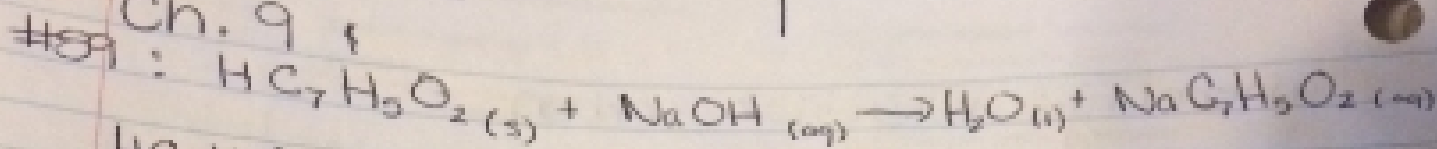


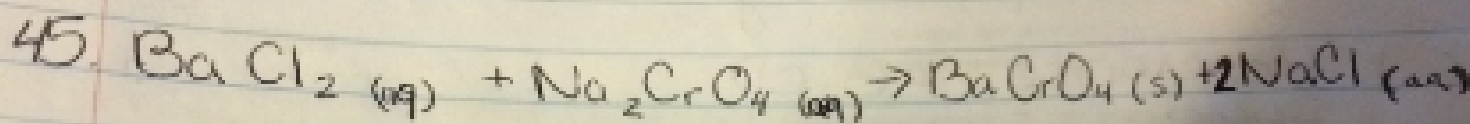
\downarrow increases atomic radius
 \uparrow increases ionization energy
 \uparrow increases electronegativity

Ch. 9



Weak acids stay together
strong acids - fall apart

Ch. 10



$1.63\text{g} \times \frac{1 \text{ mol}}{208.2\text{g}} = 7.83 \times 10^{-3} \text{ mol BaCl}_2$

$2.40\text{g} \times \frac{1 \text{ mol}}{161.98\text{g}} = 1.48 \times 10^{-2} \text{ mol Na}_2\text{CrO}_4$

$\text{BaCl}_2 \rightarrow \text{limited}$ $1.48 \times 10^{-2} \text{ BaCl}_2$
 $7.83 \times 10^{-3} \text{ Na}_2\text{CrO}_4$

$(7.83 \times 10^{-3} - 1.48 \times 10^{-2}) = 0.00697 \text{ mol excess}$

$7.83 \times 10^{-3} \left(\frac{1}{1}\right) \left(\frac{253.3\text{g}}{1}\right) = 1.98\text{g BaCrO}_4$

$0.00697 \times \frac{161.98\text{g}}{1} = 1.13\text{g excess}$

