

a) .

4) (20 points) Consider the titration curves for 25.00 mL of 0.1000 M malonic acid with 0.1000 M NaOH. For malonic acid, $K_{a1} = 1.42 \times 10^{-3}$ and $K_{a2} = 2.01 \times 10^{-6}$

e) Calculate the pH before any base is added.

f) Find the pH when 37.50 mL of base is added.

Name _____

SID _____

Score _____

_____ Points on this page

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