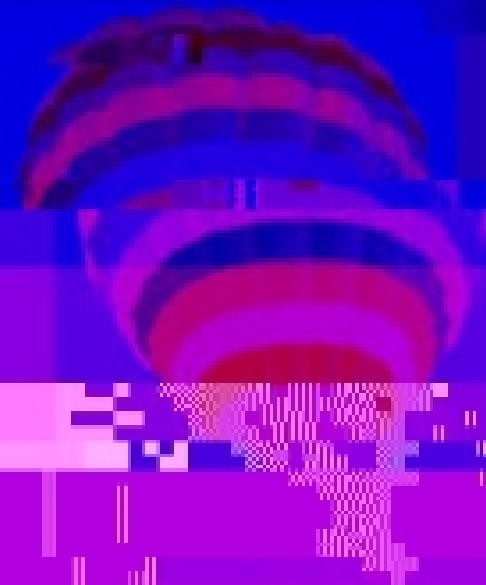


Chapter 5: Gases

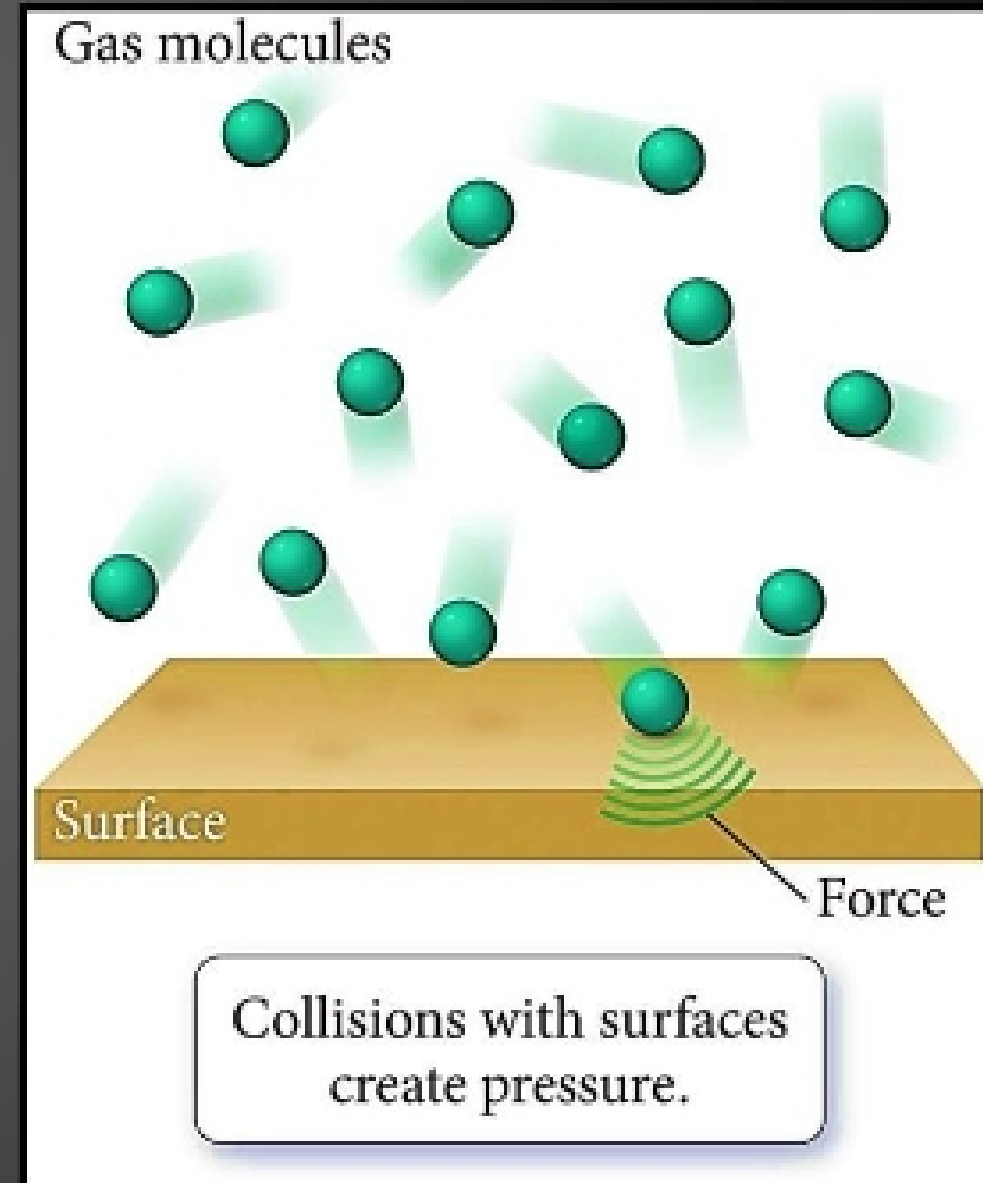


Pressure

- In gases the molecules are very far apart from each other with very little or no attraction between them
- During their motion, they collide with each other causing molecular collisions

Pressure is the *force* that results from molecular collision divided by the *area* of the surface with which they collide

$$\text{Pressure} = \frac{\text{Force}}{\text{Area}}$$



Pressure & Density

- Pressure exerted by a gas is dependent on the number of gas particles in a given volume.
- A low density of gas particles results in low pressure. A high density of gas particles results in high pressure.

