

Chem

$$\text{Ch 1: } 4.302 - 0.0122 = 4.2898$$

* S & O₂ divide by 32 & 32

Ch 4:

HCl	LiOH
HBr	NaOH
HI	KOH
	RbOH
H ₂ SO ₄	CSOH
HNO ₃	
HClO ₃	Ca(OH) ₂
HClO ₄	Sr(OH) ₂
	Ba(OH) ₂

* Moles of solute before dilⁿ = moles of solute after dilⁿ

$$* M_{\text{conc}} V_{\text{conc}} = M_{\text{dil}} \times V_{\text{dil}}$$

* Weak: Mg(OH)₂ Strong: NH₄Cl
electro. Ba(OH)₂
NaHCO₃
NH₃

* ClO₄ Perchlorate
O₃ chlorate
O₂ chlorite
O₁ hypochlorite
→ HClO₃ Chloric acid

* CO₃²⁻ carbonate
CO₂²⁻ carbonite
PO₄³⁻ - perphosphate
PO₄³⁻ phosphate
PO₃³⁻ hypophosphite

* Molecular
Ionic
pre & suffix
(III) or (II) Ammonium carbonate

Chem

- * Alkanes: C_nH_{2n+2}
- * 1 digit after ²³ d.p. CH_3
- * $N_A = 6.022 \times 10^{23} \text{ mol}^{-1}$
- * 1 mol of $CO_2 = 44 \text{ g of } CO_2$
 $= 44 \text{ amu } CO_2 \text{ (atomic wt.)}$
- * 1 mol of nitrogen = 1 mol of N_2 (not N)

↓
most common
- * Mole for = Empir. for $\times$ whole no.
- * $\% = \frac{\text{Actual}}{\text{Theor.}} \times 100$
- * $(0.0041 \times 28000.0) + (2819 \times 13)$
 $= 114.8 + (36647)$
 $= 36761.8$
 $= 3.7 \times 10^4$ (Ans should have 2 sig figs
& upto 1 decimal place)
- * $^{\circ}F = \frac{9(^{\circ}C)}{5} + 32$
- * Divide ans C_8H_8 8 by $\times$ $C_9H_{10} \rightarrow$ empirical formula
- * $44.981 = 45.0$ (3 sig fig)
- * H is active non-metal

* Smallest increment corresponds to charge on single e^-

* $H_2^{16}O$ greatest abundance

* $C_2H_{1.5} = CH_{1.49} = C_2H_3$

* NO_3^- NO_2^- NO^-
Nitrate Hypo "

* CO_3^{2-} CO_2^{2-} CO^{2-}
carbonite hypo "

* SO_4^{2-} SO_3^{2-} SO_2^{2-}
sulfite hypo

* PO_4^{3-} PO_3^{3-} PO_2^{3-}
phosphite hypo

* Dalton's model - scientific theory explaining why matters

* Same density: lump of coal in air vs water

* Heavier - less deflection

lighter - more deflection

different no of $e^- \Rightarrow$ diffⁿ deflection coz of fields

