



Determine the solubility of oxygen in water at 25 °C exposed to air at 1.0 atm. Assume a partial pressure for oxygen of 0.21 atm.



A pressure of 3.5 atm in CO_2 is required to maintain a 0.12 M CO_2 concentration in soda. Calculate the Henry's Law constant for CO_2 .



Calculate the vapor pressure at 25 °C of a solution containing 55.3 g ethylene glycol ($\text{HOCH}_2\text{CH}_2\text{CH}_2\text{OH}$) and 285.2 g water. The vapor pressure of pure water at 25 °C is 23.8 torr.